Amendments to the Specification:

Please replace the Table appearing at page 6, line 19 with the following amended Table:

Location	Reaction	Reaction Number	Delta G_f
Anode:	$2NaOH + C + H_20 \leftrightarrows Na_2CO_3 + 4H^+ + 4e^-$	1	[[+]] $\pm$0.17V
Cathode:	$4e^- + 4H^+ \leftrightarrows 2H_2$	2	
	or $4e^- + 4H^+ + O_2 \leftrightarrows 2H_2O$	3	[[+]] $\pm$1.15V
		Cell total =	[[+]] $\pm$1.32V

Please replace the Table appearing at page 6, line 21 with the following amended Table:

Location	Reaction	Reaction Number	Delta G_f
Anode:	$2KOH + C + H_20 \leftrightarrows K_2CO_3 + 4H^+ + 4e^-$	4	[[+]] $\pm$0.21V
Cathode:	$4e^- + 4H^+ \leftrightarrows 2H_2$	2	
	or $4e^- + 4H^+ + O_2 \leftrightarrows 2H_2O$	3	[[+]] $\pm$1.15V
		Cell total =	[[+]] $\pm$1.36V

Please replace the Table appearing at page 7, line 18 with the following amended Table:

Location	Reaction	Reaction Number	Delta G_f
Anode:	$Na_2CO_3 + CO_2 + H_2O \leftrightarrows 2NaHCO_3$	5	
	$NaOH + CO_2 \leftrightarrows NaHCO_3$	6	
	$3NaHCO_3 + 1.5C$ (NaCl media) $\leftrightarrows$ $3Na^+ + 3H^+ + 4.5CO_2 + 6e^-$	7	[[+]] $\pm$0.1V
Cathode:	$6e^- + 3Na^+ + 3H^+ + 1.5O_2 \leftrightarrows 3NaOH$	8	[[+]] $\pm$0.55V
		Cell total =	[[+]] $\pm$0.65V

Please replace the Table appearing at page 7, line 21 with the following amended Table:

Location	Reaction	Reaction Number	Delta G_f
Anode:	$K_2CO_3 + CO_2 + H_2O \leftrightarrows 2KHCO_3$	9	
	$KOH + CO_2 \leftrightarrows KHCO_3$	10	
	and $3KHCO_3 + C$ (KCl media) $\leftrightarrows$ $3K^+ + 3H^+ + 4.5CO_2 + 6e^-$	11	[[+]] $\pm$0.16V
Cathode:	$6e^- + 3K^+ + 3H^+ + 1.5O_2 \leftrightarrows 3KOH$	12	[[+]] $\pm$0.41V
		Cell total =	[[+]] $\pm$0.57V

Please replace the Table appearing at page 10, line 3 with the following amended Table:

Location	Reaction	Reaction Number
Anode:	$MgCO_3 + 2C + 3H_20 \leftrightarrows 3CO_2 + Mg^{2+} + 6H^+ + 8e^-$	13
Cathode:	$Mg^{2+} + 6H^+ + 8e^- + CO_2 + 2O_2 \leftrightarrows MgCO_3 + 3H_2O$	14
Net Rxn:	$2C + 2O_2 \leftrightarrows 2CO_2$ Delta G_f = [[+]] $\pm$1.02V	15

Please replace the Table appearing at page 10, line 18 with the following amended Table:

Location	Reaction	Reaction Number
Anode:	$Mg(OH)_2 + 2CO_2 \leftrightarrows Mg(HCO_3)_2$	16
	or $MgO + H_2O + 2CO_2 \leftrightarrows Mg(HCO_3)_2$	17
	$Mg(HCO_3)_2 + C \leftrightarrows Mg^{2+} + 3CO_2 + 2H+ + 4e^-$	18
Cathode:	$Mg^{2+} + 2H^+ + 4e^- + CO_2 + O_2 \leftrightarrows MgCO_3 + H_2O$	19
Net Rxn:	$2C + 2O_2 \leftrightarrows 2CO_2$ Delta G_f= [[+]] $\pm$1.02V	15

Please replace the paragraph appearing at page 8, lines 1-9 with the following amended paragraph:

These reactions of the regeneration cell, as summarized above, assume a temperature of about 104°C. The Gibbs free energy voltage for the fuel cell coupled to the regeneration cell is [[+]] =1.97 Volts for the sodium hydroxide cells and [[+]] =1.93 Volts for the potassium hydroxide cells. The voltage per mole of carbon reacted (0.92Volts for sodium and 0.89Volts for potassium) is slightly lower than the [[+]] =1.02Volts associated with carbon dioxide production. Additionally, in practice, each cell has an internal resistance and the two-cell configuration aggravates this loss. This can be partially overcome by combining the two cells in a bipolar configuration wherein a graphite-carbon electrode is modified as an anode on one side and as a cathode on the other side.